

NAME _____ DATE _____ SECTION _____

INSTRUCTOR _____ GRADE _____

EXPERIMENT 10: REPORT FOR THE TITRATIONS OF ACIDS AND BASES**DATA/RESULTS****A. STANDARDIZATION OF NaOH SOLUTION**

	example	trial 1	trial 2	trial 3
*1. Weight of KHP sample(g)	<u>0.483</u>	_____	_____	_____
*2. Volume of NaOH required (mL)	<u>21.73</u>	_____	_____	_____
3. Normality of NaOH solution (eq/L)	<u>0.109</u>	_____	_____	_____
4. Average normality of NaOH solution (eq/L)		_____		

CALCULATIONS**B. THE PERCENTAGE OF ACETIC ACID IN VINEGAR**

	example	trial 1	trial 2
*1. Volume of vinegar (mL)	<u>5.00</u>	_____	_____
*2. Volume of NaOH solution required (mL)	<u>38.23</u>	_____	_____
3. Weight of acetic acid in sample (g)	<u>0.250</u>	_____	_____
4. % of acetic acid in sample	<u>5.00</u>	_____	_____
5. Average % of acetic acid in vinegar samples		_____	

CALCULATIONS

*Numbers (items) with asterisks represent data taken in the lab, while the other numbers (items) were calculated from the lab data.

C. ANALYSIS OF ASPIRIN

	example	trial 1	trial 2
*1. Weight of aspirin tablet (g)	<u>0.896</u>	_____	_____
*2. Volume of NaOH solution required (mL)	<u>38.21</u>	_____	_____
3. Weight of ASA in tablet (g)	<u>0.750</u>	_____	_____
4. % ASA in tablet	<u>83.7</u>	_____	_____
5. Average % ASA in tablets		_____	_____

CALCULATIONS**D. THE EQUIVALENT WEIGHT OF AN UNKNOWN ACID**

	example	trial 1	trial 2	trial 3
*1. Sample number _____				
*2. Weight of unknown acid sample (g)	<u>0.196</u>	_____	_____	_____
*3. Volume of NaOH solution required (mL)	<u>24.78</u>	_____	_____	_____
4. Equivalent weight of unknown acid (g/eq)	<u>72.6</u>	_____	_____	_____
5. Average value for equivalent weight (g/eq)			_____	

CALCULATIONS

*Numbers (items) with asterisks represent data taken in the lab, while the other numbers (items) were calculated from the lab data.