



MDC Homestead Campus
Arts & Sciences Department

CHEMISTRY 1045 LABORATORY

SYLLABUS

General Information

CHM 1045L is the laboratory portion of CHM 1045. Lecture (1045) and Laboratory (1045L) are co-requisites; students must either be enrolled in both courses concurrently or have already successfully completed the lecture portion of the course. Students are required to keep a laboratory notebook. All laboratory notebook entries must strictly conform to the format set forth by your instructor. This format will be explained during the first lab meeting. Lab notebooks will be checked periodically by the instructor.

Text Book

Lab Experiments: General Chemistry, 11th ed., '06, by Pappas

Students are required to **read** the assigned laboratory exercise(s) in the lab manual **before** attending class.

Attendance and Tardiness

Attendance is REQUIRED. Any student who misses a scheduled laboratory session will be required to withdraw from the course. If the student does not withdraw he/she will automatically fail the course.

Any student arriving to class more than five minutes late will not be allowed to enter and will be considered absent.

Make-up Policy

There will be NO MAKE-UP EXAMS, except under extenuating circumstances as determined by the instructor; documentation is required. It is the student's responsibility to contact the instructor in this regard. There will be absolutely NO MAKE-UP LABS.

Scientific Notebook

Students must keep a scientific notebook (carbonless copy style, purchased in bookstore). Instructions on how to keep a scientific notebook are found in *Appendix I* of the lab manual. The scientific notebook grade will be based on the quality of the notebook's contents.

Lab Report

One typed, formal lab report must be submitted for this course. The lab report should:

- be accurate, correct, complete neat, and typed
- include calculations and setups
- use significant figures
- be turned in on time

Performance Evaluation

Students will be assigned a grade based on their performance in the lab. The instructor's grade will be based on the student's ability to follow instructions. Students are expected to:



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- follow all safety precautions (*General Safety Rules* can be found on pages vii – x in the lab manual)
- always wear safety goggles
- arrive to lab on time
- read the assigned laboratory exercise(s) in the lab manual before attending class
- clean and return all glassware/equipment to its proper place at the end of the experiment
- clean your work area

Grading Policy

Final course grades will be determined based upon the student's performance on the instructor's **performance evaluation, quizzes, weekly scientific notebook** entries, one **formal lab report** (typed), and **exams**. The point distribution will be as follows:

Task	Points
Performance Evaluation	50
Quizzes	
Quiz 1 (determined by instructor)	25
Quiz 2 (determined by instructor)	25
Quiz 3 (determined by instructor)	25
Quiz 4 (determined by instructor)	25
Scientific Notebook	
Physical & Chemical Properties (Exp. 1)	20
Densities of Liquids and Solids (Exp. 2)	20
Study of Hydrates (Exp. 4)	20
Stoichiometry (Exp. 8)	20
Analysis of Copper in a Brass Sample (Exp. 5)	20
Electrolytes and Single and Double Replacement Rxns (Appn. XIV)	20
Titrations of Acids and Bases (Exp. 10)	20
The Ideal Gas Equation (Exp. 7)	20
Atomic Spectra (Exp. 6)	20
Molecular Weight Determination of a Pure Substance (Exp. 9)	20
Formal Lab Report	100
Exams	
Exam 1 (Exp. 1, 2, 4, 8)	100
Exam 2 (Exp. 5, 10, Appn. XIV)	100
Exam 3 (Exp. 6, 7, 9)	100
Total Points	750

Letter Grade	Percent	Points Needed
A	88% - 100%	660
B	78% - 87%	585
C	65% - 77%	488
D	57% - 64%	428
F	0% - 56%	less than 428

Academic Dishonesty Policy:

See Miami Dade College Students Rights and Responsibilities PROCEDURE 4035.



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SCHEDULE

Week	Topic	Experiment
1	Course Intro. & Lab Safety	
2	Physical & Chemical Properties Parts A & D Instructor Demonstration	1
3	Densities of Liquids and Solids	2
4	Study of Hydrates	4
5	Stoichiometry	8
6	EXAMINATION I	
7	Analysis of Copper in a Brass Sample	5
8	Electrolytes and Single and Double Replacement Rxns	Appendix XIV
9	Titrations of Acids and Bases	10
10	Titrations of Acids and Bases (continued)	10
11	EXAMINATION II	
12	The Ideal Gas Equation	7
13	Atomic Spectra	6
14	Molecular Weight Determination of a Pure Substance	9
15	EXAMINATION III	